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Supplement of

Chlorine-initiated oxidation of *n*-alkanes under high-NO_x conditions: insights into secondary organic aerosol composition and volatility using a FIGAERO-CIMS

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S1 Additional Table and Figures

Table S1. SOA Bulk Composition and Alternative Yields

Exp #	Precursors	RH	Y _{SOA} ^a	MW ^b	Y _{SOA,m} ^c	ρ ^d	Y _{SOA,L} ^e	f _{HClA} ^f	f _{ClF} ^g	f _{OCl} ^h	f _{ON} ⁱ	R _{ON} ^j
1	Octane, NO, Cl ₂	5 >	0.28	226	0.14	2.3	0.13	1.3 E ⁻²	1.8 E ⁻²	0.15	0.54	0.64
2	Octane, NO ₂ , Cl ₂	5 >	0.16	223	0.08	1.7	0.10	1.1 E ⁻²	1.7 E ⁻²	0.14	0.49	0.57
3	Octane, NO, NO ₂ , Cl ₂	5 >	0.24	206	0.13	1.8	0.14	1.0 E ⁻²	2.1 E ⁻²	0.15	0.49	0.57
4	Octane, NO, NO ₂ , Cl ₂	35	0.24	223	0.12	2.1	0.12	1.4 E ⁻²	1.9 E ⁻²	0.14	0.49	0.57
5	Decane, NO, Cl ₂	5 >	0.84	265	0.45	2.3	0.39	1.1 E ⁻²	1.5 E ⁻²	0.13	0.60	0.84
6	Decane, NO ₂ , Cl ₂	5 >	0.45	249	0.26	2.0	0.24	1.0 E ⁻²	1.6 E ⁻²	0.14	0.56	0.75
7	Decane, NO, NO ₂ , Cl ₂	5 >	0.80	241	0.47	2.1	0.40	1.2 E ⁻²	1.6 E ⁻²	0.12	0.59	0.75
8	Decane, NO, NO ₂ , Cl ₂	40	0.50	243	0.29	1.6	0.33	1.2 E ⁻²	1.7 E ⁻²	0.13	0.56	0.71
9	Dodecane, NO, Cl ₂	5 >	1.65	270	1.04	1.8	0.99	0.8 E ⁻²	1.7 E ⁻²	0.14	0.67	0.93
10	Dodecane, NO ₂ , Cl ₂	5 >	1.25	258	0.82	2.1	0.63	1.4 E ⁻²	1.7 E ⁻²	0.14	0.60	0.78
11	Dodecane, NO, NO ₂ , Cl ₂	5 >	1.40	264	0.90	2.4	0.62	1.4 E ⁻²	1.8 E ⁻²	0.15	0.62	0.82
12	Dodecane, NO, NO ₂ , Cl ₂	67	1.10	260	0.72	2.4	0.49	1.1 E ⁻²	1.6 E ⁻²	0.13	0.61	0.77

- (a) Assuming a CE of 0.5 and RIE of 1.4 for ACSM, in $\mu\text{g m}^{-3}$
- 5 (b) Average molecular weight calculated based on organic ions (I⁻ adducts only) observed by FIGAERO-CIMS in g mol^{-1}
- (c) SOA molar yield, calculated using Eqs. S1-3, in mol mol^{-1}
- (d) Density calculated based ACSM mass measurements and SEMS volume measurements
- (e) SOA mass yield, recalculated assuming the lowest reported alkane-OH density, 1.06 g cm^{-3} (Lim and Ziemann, 2009), in $\mu\text{g m}^{-3}$
- 10 (f) Mass fraction of -Cl functional group, estimated using the ACSM measurement of HCl⁺ ion, relative to the bulk SOA
- (g) Mass fraction of -Cl functional group, estimated using the SOA molecular composition as measured by the FIGAERO-CIMS, relative to the total organic ions observed
- (h) Mass fraction of organochlorides (including chloronitrates), estimated using the SOA molecular composition as measured by the FIGAERO-CIMS, relative to the total organic ions observed
- 15 (i) Mass fraction of organonitrates (excluding chloronitrates), estimated using the SOA molecular composition as measured by the FIGAERO-CIMS, relative to the total organic ions observed
- (j) Average ratio of -NO₃ functional group to the number of carbon, estimated using the SOA molecular composition as measured by the FIGAERO-CIMS. Individual -NO₃ functional groups in dinitrates and trinitrates are counted separately (e.g. a C₁₂ dinitrate has the same R_{ON} as a C₆ mononitrate).

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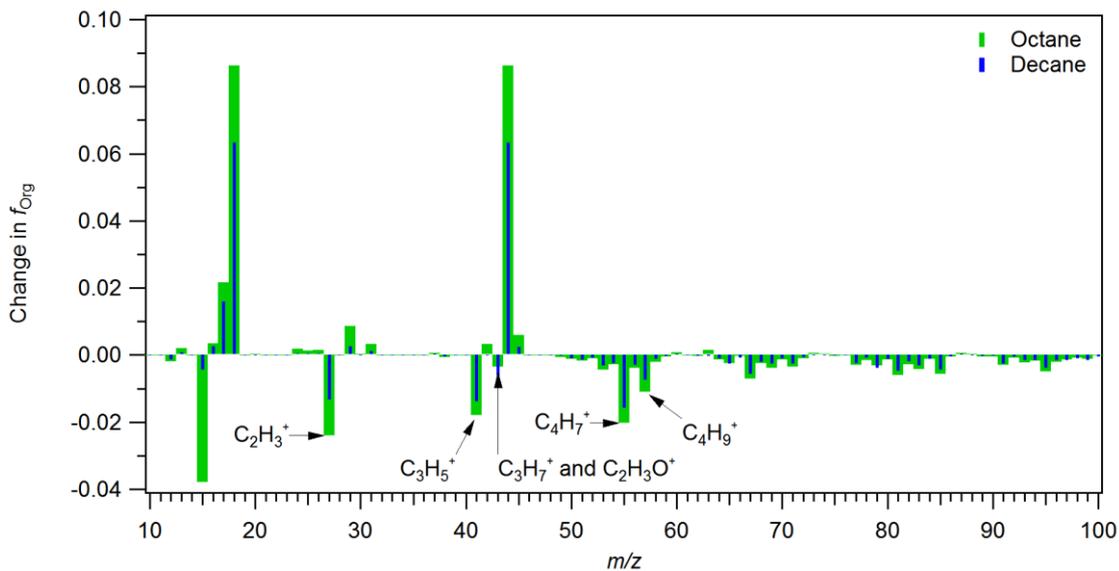


Figure S1: Change in the fractional contribution by different organic ion fragments to the overall organic aerosol mass as measured by the ACSM. Dodecane-Cl SOA (Exp. 11) was used as the reference for octane-Cl SOA (Exp. 3) and decane-Cl SOA (Exp. 7) produced under similar oxidation conditions under low RH. Mass spectra shown are 50-minute averages from minute 10 to 60 during the photooxidation period. Ion identities indicated ($C_2H_3^+$, $C_3H_5^+$, $C_3H_7^+$, $C_2H_3O^+$, $C_4H_7^+$, $C_4H_9^+$) were assumed, as the quadruple ACSM used lacks the necessary mass resolution for exact identification. At m/z 43, the $C_2H_3O^+$ is expected to dominate. Highlighted C_xH_y ions are often associated with hydrocarbon-like organic aerosol (Ng et al., 2011; Ulbrich et al., 2009). Clear enhancement of m/z 44 and associated ion fragments (e.g. m/z 16, 17, and 18) was also observed for octane and decane SOA.

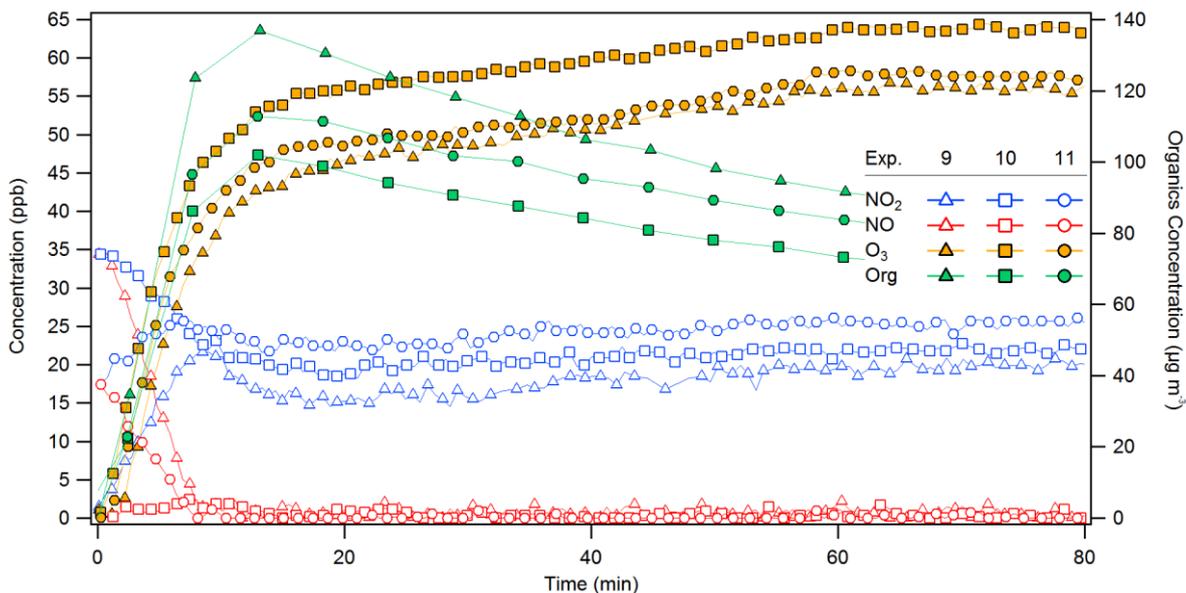


Figure S2: Example of NO, NO₂, O₃ and SOA trends under different starting NO_x conditions. Exp. 9 (high initial NO), 10 (high initial NO₂), and 11 (balanced initial NO and NO₂) for chlorine-initiated oxidation of dodecane under low RH conditions are shown. Ozone production slowed down significantly as the NO_x concentrations plateaued and as the SOA concentrations approached their maxima. By the end of the photooxidation period (60 min), lowest final NO_x concentrations (consisting of NO₂ and interferences, which may include alkyl nitrates) and highest SOA concentrations were observed for high initial NO concentrations (Exp. 1 for octane, Exp. 5 for decane, and Exp. 9 for dodecane).

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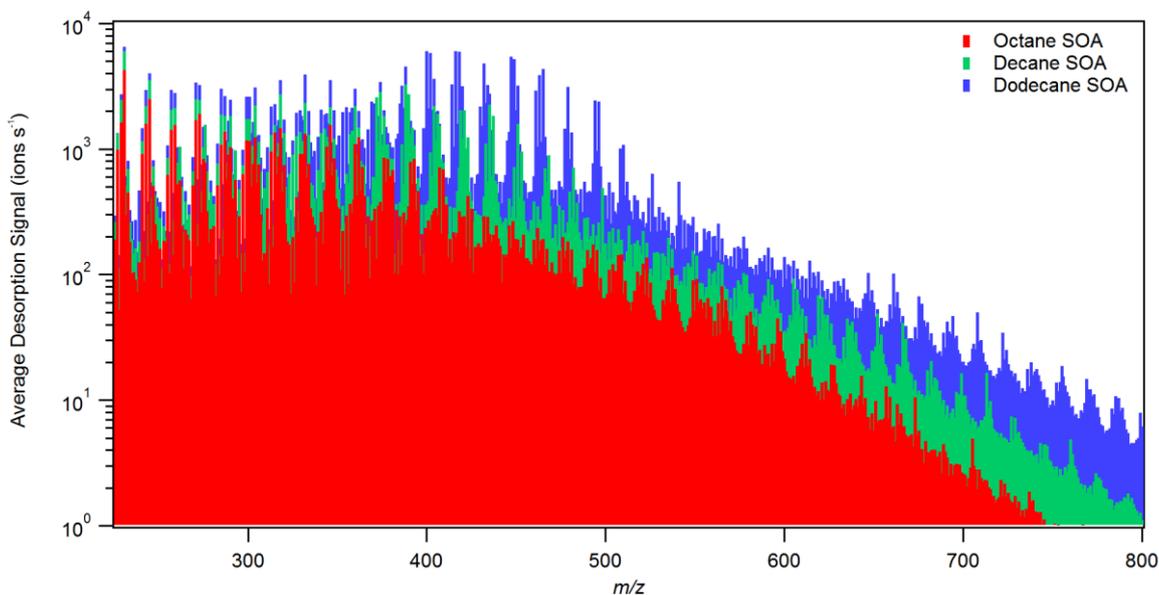
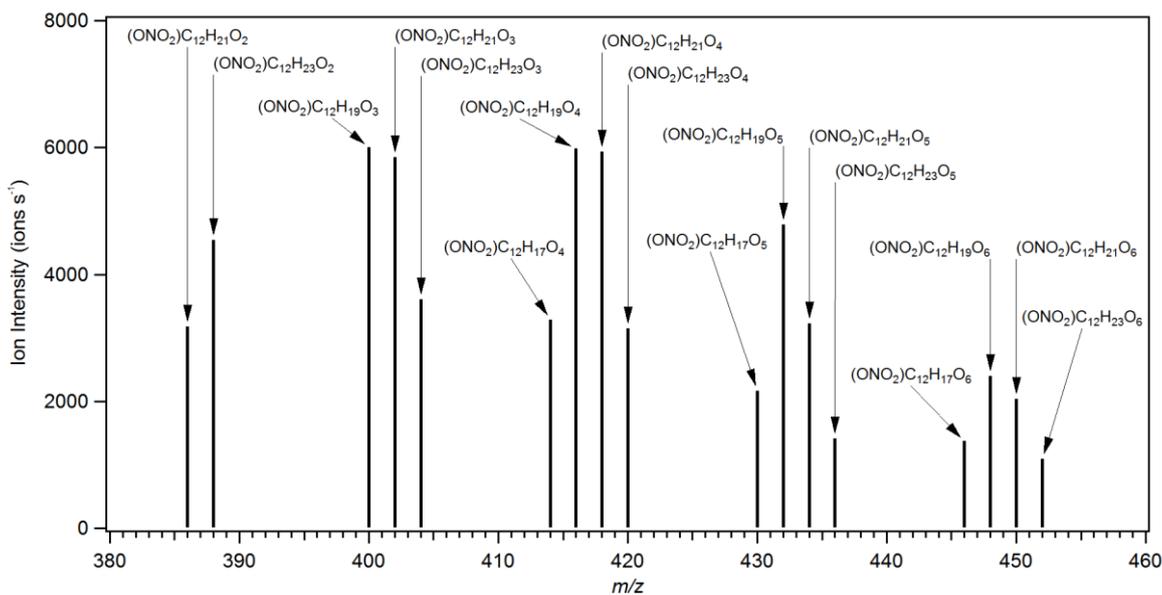


Figure S3: Comparison of particle-phase composition during FIGAERO desorption for octane (Exp. 3), decane (Exp. 7), and dodecane (Exp. 11). Same as Figure 4 in the main text but with the y-axis in logarithmic scale and the x-axis (m/z) extended to the maximum sampled range. Desorption signals are not stacked.



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Figure S4: Particle-phase distribution of C_{12} organic (mono-)nitrates. Assuming equal sensitivity, the particle-phase abundance roughly followed a bell-shape distribution across the different oxygenation groups, peaking at the O_3 or O_4 group. Within each oxygenation group, the product distribution follows a bell-shape around H_{19} and H_{21} compounds. Similar bell-shaped organic nitrate distribution has been observed for ambient isoprene- and monoterpene-derived organonitrates (Lee et al., 2016).

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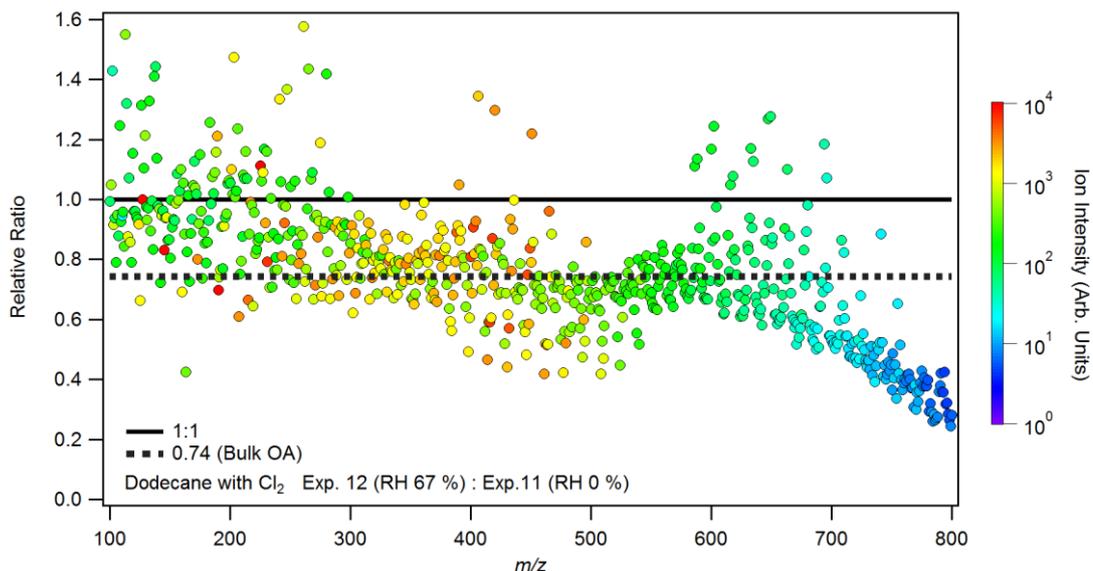


Figure S5: Ratio of integrated unit-mass resolution FIGAERO-CIMS desorption signal during temperature ramp from a high-RH SOA formation experiment (Exp. 12, 67 % RH) to that from a low-RH SOA formation experiment (Exp. 11, < 5 % RH). Integrated ion intensity is shown on a logarithmic scale as the color. The ratio of bulk organics concentration during the filter collection period calculated using ACSM data is shown as the dotted black line. Except for the low molecular weight thermal fragmentation/decomposition products and the high molecular weight, low-volatility compounds, the unit-mass ion ratios appear in agreement with bulk measurement. Under high RH conditions, the higher molecular weight ($m/z > 600$) compounds become increasingly suppressed.

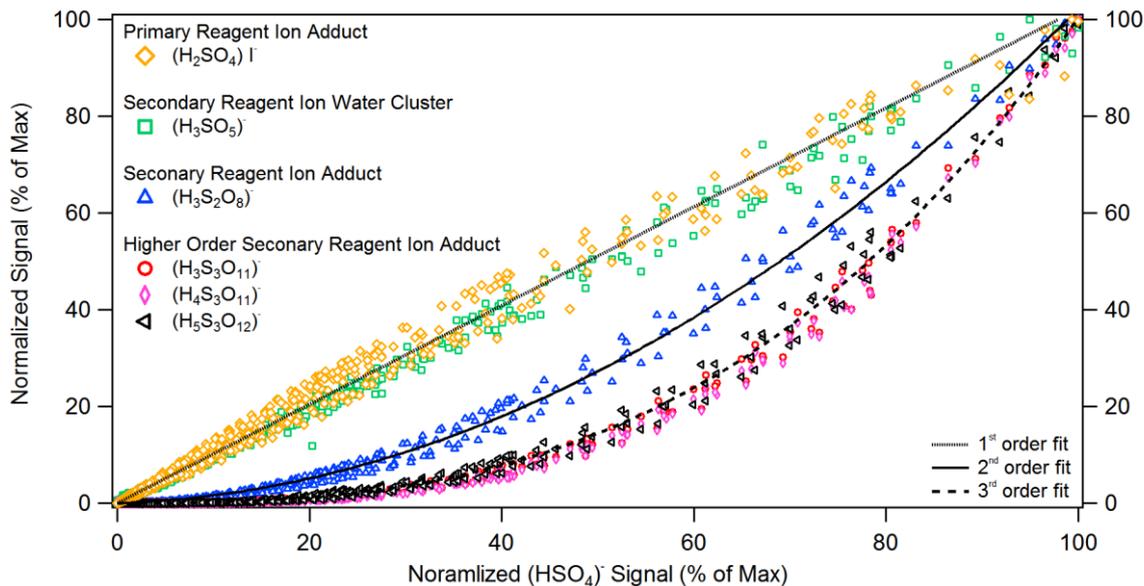


Figure S6: Example of secondary HSO_4^- ionization products. Ion signals observed during $(\text{NH}_4)_2\text{SO}_4$ thermal decomposition were first normalized by I^- signal and then normalized against the maximum desorption signal observed for each ion. Linear (1st order), quadratic (2nd order), and cubic equations (3rd order) are fitted.

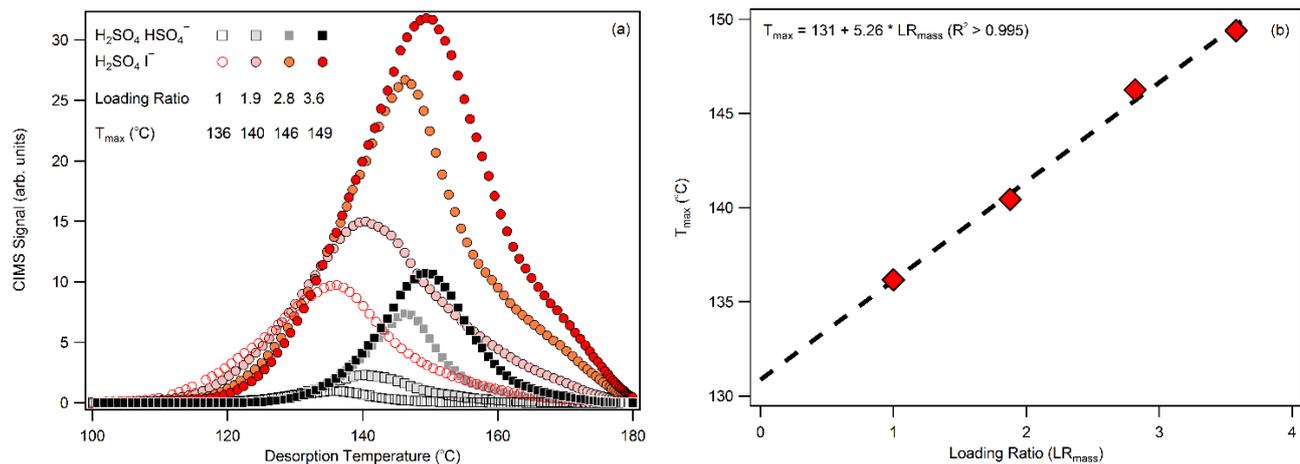


Figure S7: The 1-D thermograms for two (NH₄)₂SO₄ decomposition product ions. H₂SO₄•I⁻ was the primary adduct and H₂SO₄•HSO₄⁻ was a secondary ionization product. The lowest time-integrated desorption ion intensity for H₂SO₄•I⁻ was used as the basis for calculating the loading ratio (LR_{mass}). The loading ratio for H₂SO₄•HSO₄⁻ (not shown) is the square of that for H₂SO₄•I⁻ (b) Linear regression fitting of T_{max} as a function of the LR_{mass}. T_{max} increased linearly with mass loading of (NH₄)₂SO₄.

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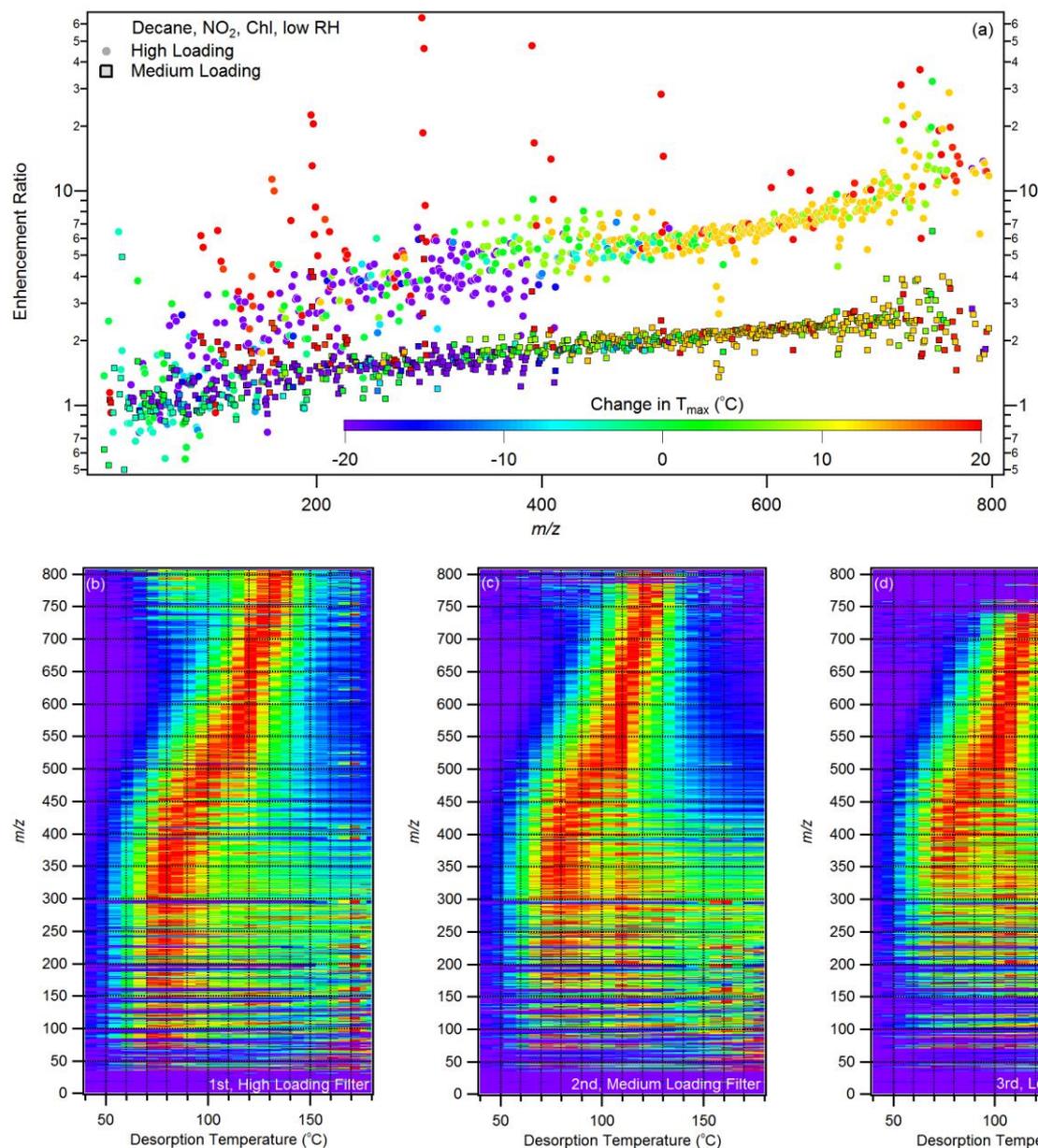


Figure S8: Effects of aerosol loading on T_{max} . Three SOA filter samples were collected during Exp. 6 for 45 mins, 30 mins, and 15 mins, in that order. Each filter was subjected to two full desorption runs to minimize carry-over effects and to determine the filter background.

- 5 Enhancement ratios calculated using unit-mass integrated signals for the lowest filter loading run as the referenced condition. Color scale shows T_{max} shifts in °C, using the T_{max} values observed in the low loading run as the reference point. (b) 2-D thermogram for the first, high loading filter, (c) for the second, medium loading filter, and for (d) the third, low loading filter. For (b-d), the color scale represents the normalized desorption intensity as a percentage of the maximum, same as in Fig. 5 in the main text. Between the time it took for filter collection and filter desorption, the ion intensity of some high molecular weight, low volatility compounds had decreased to below the limits of detection (i.e. 3σ of the background signal) in the particle phase, likely due to volatility-dependent wall loss. Disappearance of low m/z (< 300) compounds maybe due to decreases in suspended aerosol concentration, making it unfavorable for semi-volatile compounds to partition to the particle phase, or may be associated with the loss of high molecular weight oligomers, assuming that the low m/z desorption ions were dominated by low-temperature ($T_{max} < \sim 80$ °C) thermal fragmentation products as opposed to semi-volatile monomers.
- 10

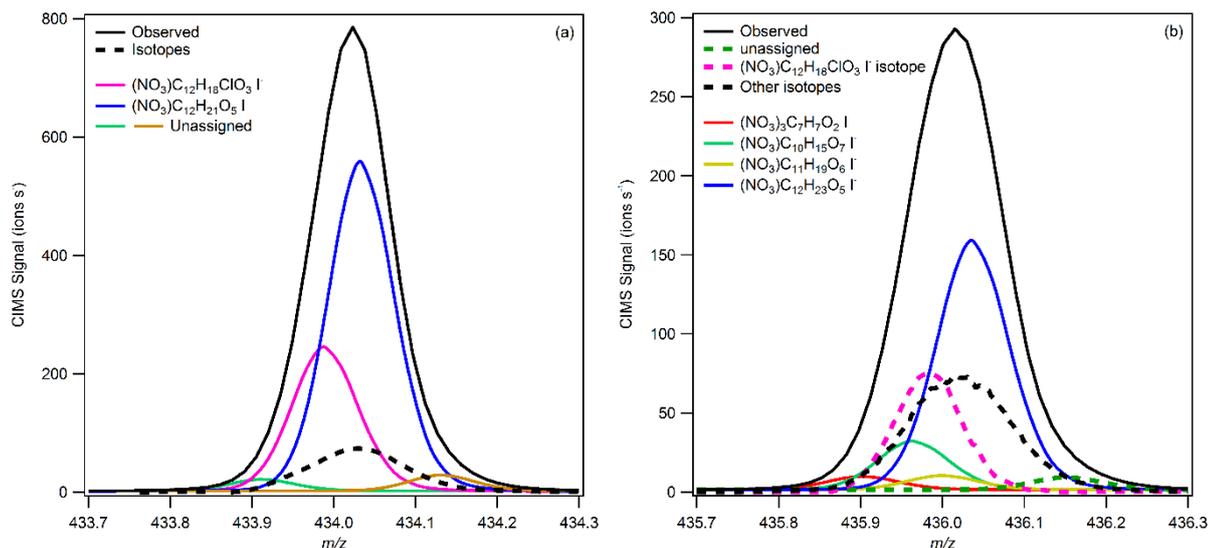


Figure S9: High resolution fitting at m/z (a) 434 and (b) 436 for FIGAERO-CIMS data from Exp. 11. The C_{12} chloronitrate ($ONO_2-C_{12}H_{18}ClO_3 \cdot I$) peak is tentatively identified in (a) but it overlaps with the nearby stronger organonitrate ($ONO_2-C_{12}H_{21}O_5 \cdot I$) peak. Because the chloronitrate peak is a weaker peak with significant overlap with a stronger peak, quantitative assessment would be challenging due to peak fitting uncertainties (Cubison and Jimenez, 2015). The presence of chloronitrate cannot be confirmed or rejected based on its unique isotopic signature at $m/z+2$ positions, as shown in (b), where nonchlorinated compounds also dominate.

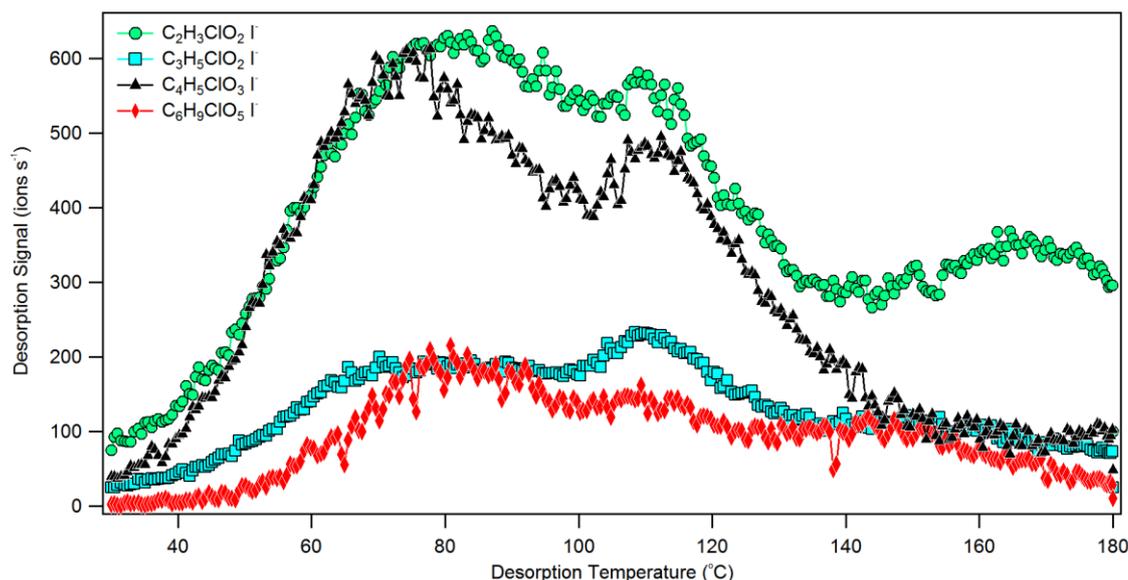


Figure S10: 1-D thermogram of select organochlorides observed for Exp 10. All compounds shown exhibited multimodal desorption behaviors. $C_2H_3ClO_2$ is too volatile to be present as a molecular compound in the particle phase and is therefore a thermal decomposition product. $C_2H_3ClO_2$ shows three local maxima at ~ 87 °C, ~ 109 °C, and ~ 167 °C. The T_{max} for the least volatile desorption mode of $C_2H_3ClO_2$ is higher than that for ammonium sulfate seed particles (~ 155 °C) from the same filter run, and may be produced from thermal decomposition of extremely low volatility organochlorides. However, larger organochloride ions were not observed to have any distinct T_{max} modes over these very high temperature ranges (> 160 °C)

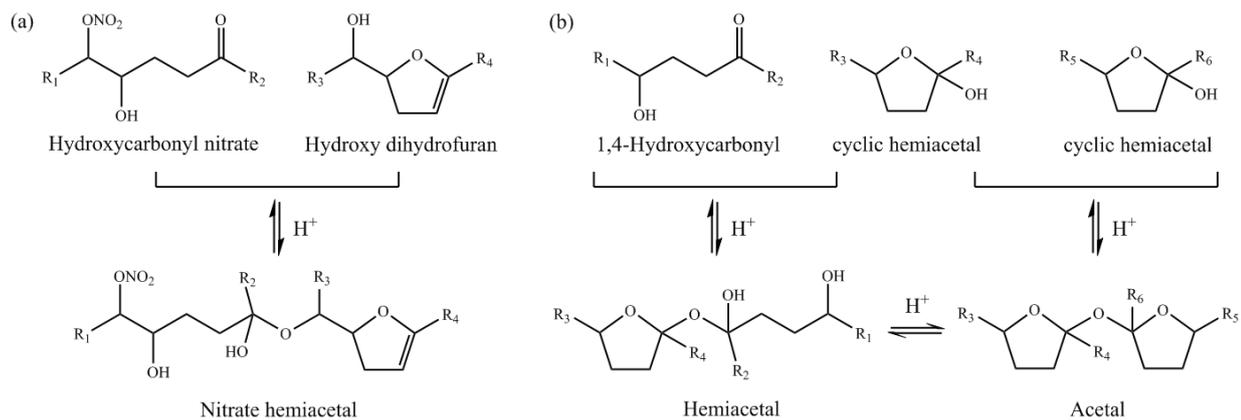


Figure S11: Oligomerization via the reaction between (a) carbonyl hydroxyl nitrate and hydroxy dihydrofuran proposed by Schilling Fahnstock et al. (2015) and (b) cyclic hemiacetal with either a 1,4-hydroxycarbonyl or another cyclic hemiacetal proposed by Aimanant and Ziemann (2013). Condensed-phase isomerization between acetal and hemiacetal dimers is also possible, as shown in (b).

S2. Interaction between Cl, NO_x, and HO_x

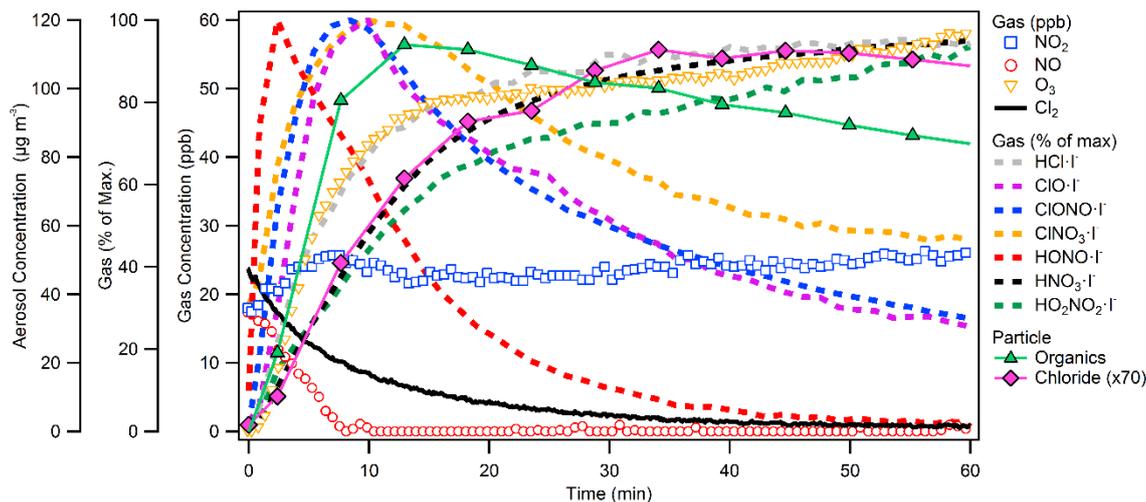
Formation of HO₂NO₂ likely proceeds via



5 Formation of HONO under UV can proceed via



The trends for HO₂NO₂, HONO, and select gas-phase species observed by the I⁻ CIMS are shown in Fig. S12.



10 **Figure S12:** Representative trends of SOA and trace gas species during the photooxidation period. Data from dodecane oxidation (Exp. 11) are shown, similar to Fig. 1 in the main text. Additional species shown include ClONO₂ (from reaction between ClONO with NO₂, Lesar et al., 2006), ClO (possibly a ion fragment of ClONO), HCl (from H-abstraction by Cl[•] or perhaps Cl-elimination reactions), and HNO₃ (due to the oxidation of NO₂ by OH).

S3. Calculation of Bulk SOA Properties using the FIGAERO-CIMS

15 If equal sensitivity is assumed for all ions (i.e. I⁻ adducts detected by the FIGAERO-CIMS), the average molecular weight, mw_{avg} is estimated based on the (desorption) intensity I_i and the molecular weight, mw_i of all ions identified,

$$mw_{avg}(T_d) = \frac{\sum_i mw_i \times I_i(T_d)}{\sum_i I_i(T_d)} \quad Eq. (S1)$$

where mw_{avg} and I_i vary with desorption temperature, T_d during a single FIGAERO desorption. A single average molecular weight for the entire FIGAERO desorption run can be calculated based on the integrated values over the T_d range,

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$$mw_{SOA} = \frac{\int (\sum_i mw_i \times I_i(T_d)) \Delta T_d}{\int (\sum_i I_i(T_d)) \Delta T_d} \quad Eq. (S2)$$

where ΔT_d is the step change in desorption temperature. It should be noted that because of thermal decomposition, which has also been observed for monomers (Stark et al., 2017), the mw_{avg} and mw_{SOA} calculated using the FIGAERO-CIMS data likely underestimate the actual average SOA molecular weight. For octane SOA, mw_{SOA} calculated in this way ranges from 206 to

226 g mol⁻¹; for decane SOA, mw_{SOA} ranges from 241 to 265 g mol⁻¹; for dodecane SOA, mw_{SOA} ranges from 260 to 270 g mol⁻¹. The molar SOA yield can then be calculated as,

$$Y_{molar} = \frac{n_{SOA}}{\Delta n_{VOC}} = \frac{m_{SOA}/mw_{SOA}}{\Delta m_{VOC}/mw_{VOC}} = Y \frac{mw_{VOC}}{mw_{SOA}} \quad \text{Eq. (S3)}$$

where n_{SOA} and n_{VOC} are the molar concentrations of SOA and VOC. The molecular weight of VOC, mw_{VOC} is known for the alkane precursors. The corresponding molar yield ranges from 0.08 to 0.14 for octane SOA, 0.26 to 0.47 for decane SOA, and 0.72 to 1.04 for dodecane SOA. The above-unity molar SOA yield observed for dodecane (Exp. 9) indicates that SOA mass may be overestimated, which may be the result of uncertainties with the collection efficiency (CE) and relative ionization efficiency (RIE) assumed for the SOA.

Equations S1 and S2 can also be used to calculate various other bulk SOA properties, including elemental ratios such as the oxygen-to-carbon ratio (O:C), hydrogen-to-carbon ratio (H:C), nitrate-to-carbon ratio (NO₃:C), chloride-to-carbon ratio (Cl:C), or the oxidation state of carbon (OS_C). Using the oxygen-to-carbon ratio (O:C) as the example, the bulk SOA elemental ratios as a function of the FIGAERO desorption temperature, T_d can be calculated as

$$O:C_{avg}(T_d) = \frac{n_O(T_d)}{n_C(T_d)} = \frac{\sum_i n_{O,i} \times I_i(T_d)}{\sum_i n_{C,i} \times I_i(T_d)} \quad \text{Eq. (S4)}$$

where n_O is the total number of oxygen atoms present in the desorbed ions at T_d, n_C is the total number of carbon atoms, n_{O,i} is the number of oxygen atom present in compound *i* as determined by its assigned molecular formula (which is independent of T_d), n_{C,i} is the number of carbon present in compound *i*, and I_i is the desorption ion intensity for compound *i* at T_d. Equal sensitivity is assume for ions used for the analysis such that I_i can be used as the molar amount for compound *i*.

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