

# Atmospheric acidification of mineral aerosols: a source of bioavailable phosphorus for the oceans

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**Abstract.** Primary productivity of continental and marine ecosystems is often limited or co-limited by phosphorus. Deposition of atmospheric aerosols provides the major external source of phosphorus to marine surface waters. However, only a fraction of deposited aerosol phosphorus is water soluble and available for uptake by phytoplankton. We propose that atmospheric acidification of aerosols is a prime mechanism producing soluble phosphorus from soil-derived minerals. Acid mobilization is expected to be pronounced where polluted and dust-laden air masses mix. Our hypothesis is supported by the soluble compositions and reconstructed pH values for atmospheric particulate matter samples collected over a 5-yr period at Finokalia, Crete. In addition, at least tenfold increase in soluble phosphorus was observed when Saharan soil and dust were acidified in laboratory experiments which simulate atmospheric conditions. Aerosol acidification links bioavailable phosphorus supply to anthropogenic and natural acidic gas emissions, and may be a key regulator of ocean biogeochemistry.

## 1 Introduction

Primary productivity of continental and marine ecosystems is often limited or co-limited by phosphorus (Howarth, 1988; Elser, 2007). Because riverine inputs of phosphorus (P) preferentially accumulate along the continental margins, airborne sources of P are particularly important for supporting primary production in oligotrophic areas of the open ocean (Paytan and McLaughlin, 2007). Similar to the nutrient elements iron and silicon, most P in the atmosphere is associated with particulate matter, primarily mineral aerosols (Mahowald et al., 2008). The major forms of P in soil-derived dust are minerals from the apatite group and, to a lesser extent, P bound to iron (hydr)oxide minerals (Singer et al., 2004; Eijsink et al., 2000). These minerals are highly insoluble under the alkaline and oxygenated conditions encountered in oceanic surface waters (Atlas and Pytkowicz, 1977). Combined with the short transit times of mineral aerosols through the photic zone, this implies that the main source of bioavailable P in atmospheric deposition is water soluble P produced during airborne processing of soil-derived dust. Mahowald et al. (2008) estimate that globally about 17 % of total atmospheric P deposited at the sea surface is water soluble. The soluble fraction, however, is highly variable, with values ranging between 7 and 100 % (Mahowald et al., 2008). Baker et al. (2006) observed a gradient in soluble P in dust advected off the coast of North Africa. Mahowald et



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al. (2008) further note that it is not known at present what processes control the amount of soluble P in aerosols.

We hypothesize that acid processing of mineral aerosols is a major pathway for the production of water soluble P in the atmosphere. The main acids in the atmosphere,  $\text{H}_2\text{SO}_4$  and  $\text{HNO}_3$ , are generated by the oxidation of sulfur and nitrogen gases emitted by biogenic, volcanic and anthropogenic sources (Seinfeld and Pandis, 2006). These acids, if present in sufficient amounts in mineral aerosols, will overcome the available carbonate buffer capacity. This causes aerosol pH to drop to low values under clear-sky conditions (Meskhidze et al., 2003), increasing the solubilities of apatite and iron (hydr)oxide minerals by several orders of magnitude (Stumm and Morgan, 1996). Production of soluble P is analogous to that invoked to explain the presence of soluble iron and trace metals in atmospheric dust (Meskhidze et al., 2003; Spokes et al., 1994).

## 2 Evidence of acid mobilization of P in Eastern Mediterranean aerosol

The Eastern Mediterranean Sea (EMS) offers an ideal setting for testing the hypothesis of soluble (leachable) inorganic P (LIP) production by aerosol acidification. The EMS is unequivocally P limited with a molar nitrate-to-phosphate ratio of 28:1 in the deep water (Krom et al., 1991). Dry deposition is a major source of P to the basin and directly affects primary productivity in offshore areas (Krom et al., 2004). The Saharan desert is the principle source region of mineral aerosols reaching the central EMS and can interact with polluted air masses from Europe and the Middle East. Size segregated aerosol samples were collected over a period of five years at the Finokalia station ( $35^\circ 20' \text{ N}$ ,  $25^\circ 40' \text{ E}$ ), a remote marine background site on Crete located 50 m inland and 230 m a.s.l. (Sciare et al., 2003). HYSPLIT back trajectories ([www.arl.noaa.gov/ready/hysplit4.html](http://www.arl.noaa.gov/ready/hysplit4.html)) indicate that the air masses sampled at Finokalia originate from the marine boundary layer, Northern Africa, Europe, Asia Minor and the Middle East. Air masses from Europe, the former Soviet Republics and Asia Minor tend to exhibit the highest anthropogenic pollutant loadings (aerosol sulfates and nitrates), while the Mediterranean and Northern Africa typically yields cleaner air masses (Mihalopoulos et al., 1997). Hence, Saharan dust reaching Finokalia has experienced variable degrees of interaction with polluted air.

### 2.1 Collection and chemical analysis of aerosol samples

Aerosol samples were collected using a virtual impactor (VI; Loo and Cork, 1988). The VI was modified to divide particles into two size fractions: fine (aerodynamic particle diameter,  $D_a < 1.3 \mu\text{m}$ ) and coarse particles ( $D_a > 1.3 \mu\text{m}$ ). The inlet preceding each VI has a cut-off size of  $10 \mu\text{m}$ . The operational flow rate is  $16.71 \text{ min}^{-1}$ , divided into 1.7 and

$15.01 \text{ min}^{-1}$  minor and major flows, respectively. The average sampling time was two days (from 1 to 3 days). Polytetrafluoroethylene (PTFE) filters (Millipore Fluoropore; pore size  $3.0 \mu\text{m}$ ; diameter 47 mm) were used for sampling.

Samples were pre- and post-weighed using a 6-digit microbalance (ATI-CAHN/CA27). The readability of the balance is  $10 \mu\text{g}$  with a precision of  $40 \mu\text{g}$  corresponding to mass concentration uncertainties of 0.86 and  $0.77 \mu\text{g m}^{-3}$  for fine and coarse fractions, respectively. After weighing, the samples were stored in a freezer in Petrislides (Millipore Inc.). Samples were analyzed for water-soluble ions and LIP. One quarter of each PTFE filter was extracted using 20 ml of nanopure water. The solutions obtained were analyzed by ion chromatography (IC) for anions ( $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{NO}_3^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{C}_2\text{O}_4^{2-}$ ) and cations ( $\text{Na}^+$ ,  $\text{NH}_4^+$ ,  $\text{K}^+$ ,  $\text{Mg}^{2+}$  and  $\text{Ca}^{2+}$ ). Accuracy was determined by comparing results with certified WMO and DIONEX standards and was better than 5 % for all elements ( $n=5$ ). More details on the IC method are given in Bardouki et al. (2003). LIP was analyzed colorimetrically following Bardouki et al. (2003).

### 2.2 Thermodynamic modelling of atmospheric samples

The chemical composition, speciation and phase state of the aerosols collected on the filter samples were modelled using the ISORROPIA/ISORROPIA-II aerosol thermodynamic models (<http://nenes.eas.gatech.edu/ISORROPIA>; Nenes et al., 1998; Fountoukis et al., 2007). These models have been extensively evaluated against in-situ observations (e.g., Yu et al., 2005; Fountoukis et al., 2009), and are currently used for driving simulations in US, Asian and European global, regional and air quality models.

ISORROPIA/ISORROPIA-II simulates the composition of aerosol in equilibrium with the surrounding gas phase, including the transformations taking place when freshly emitted dust is mixed with acidic pollution. The model takes as input the amount of “aerosol precursor” sodium, potassium, ammonium (gas-phase  $\text{NH}_3$  plus aerosol ammonium), sulfate (in the form of neutral sulfate and bisulfate ions), magnesium, calcium, chloride (gas-phase  $\text{HCl}$  plus aerosol chloride), nitrate (gas-phase  $\text{HNO}_3$  plus aerosol nitrate), relative humidity and temperature. Based on this input, the model predicts at thermodynamic equilibrium the phases present in the aerosol particles (aqueous, solid, or both), the amount and chemical composition of each phase, and, the concentrations of semi-volatile species (i.e.,  $\text{NH}_{3(\text{g})}$ ,  $\text{HNO}_{3(\text{g})}$ ,  $\text{HCl}_{(\text{g})}$ ) in the gas phase.

Compositions are calculated by solving a system of equations derived from the equilibrium reactions and chemical potential of the species involved. The number of equations and iterations required is minimized by considering compositional “regimes”; because of this, ISORROPIA-II is considered one of the most computationally efficient thermodynamic equilibrium models available. Activity coefficients are calculated using pre-calculated lookup tables or

a combination of Kusik-Meissner and Bromely models (the latter of which was used in this study). The amount of water contained within the aerosol is in equilibrium with the gas phase, so that the water activity must equal to the fractional relative humidity.

Predictions were carried out assuming that the aerosol particles are in equilibrium at an ambient relative humidity of 95 %, a temperature of 298 K, and with compositions constrained from the observations. The equilibrium assumption applies well to submicron (fine) aerosol but may introduce errors when applied to coarse mode aerosol (e.g., Capaldo et al., 2000). Thermodynamic calculations are therefore carried out for all the fine aerosol samples, and for well-aged coarse aerosol samples (for which the carbonate buffer has been neutralized). The pH of the aerosol is predicted by ISORROPIA/ISORROPIA-II; the concentration of total dissolved P is calculated by dividing the measured LIP with the calculated aerosol liquid water content.

### 2.3 Apatite solubility calculations

Dissolution of apatite,  $\text{Ca}_5(\text{PO}_4)_3(\text{OH}, \text{F}, \text{Cl})$ , is the primary source of mineral phosphorus in soils (Newman, 1995; Nezat et al, 2007). Reported solubility products for apatite minerals are highly variable, because of solid solution formation, incorporation of impurities (e.g., carbonate) and non-stoichiometric composition. Solubility calculations were carried out considering the pure end-member minerals hydroxyapatite (HAP) and the less-soluble fluoapatite (FAP) in NaCl solutions. The calculations provide a lower estimate of solubility, as the formation of aqueous phosphate complexes – in particular with magnesium – will increase the solubility of an apatite mineral relative to that in a NaCl solution of the same ionic strength (Hershey et al., 1989). The incorporation of carbonate in the apatite structure further enhances the mineral solubility relative to that of the pure end-member phases (Jahnke, 1984). Dissolution of HAP and FAP is assumed to occur in a background electrolyte solution (i.e., containing no dissolved calcium, phosphate or fluoride initially) using the following solubility products for HAP and FAP at 25 °C and 1 bar:  $K_{\text{HAP}} = (\text{Ca}^{2+})^5(\text{HPO}_4^{2-})^3(\text{OH}^-)/(\text{H}^+)^3 = 10^{-20.47}$  and  $K_{\text{FAP}} = (\text{Ca}^{2+})^5(\text{HPO}_4^{2-})^3(\text{F}^-)/(\text{H}^+)^3 = 10^{-23.12}$  (Van Cappellen and Berner, 1991), where  $(X)$  denotes the activity of ionic species  $X$  in solution. The concentration of LIP,  $m_{\text{TP}}$ , is related to that of  $\text{HPO}_4^{2-}$  via  $m_{\text{TP}} = m_{\text{HPO}_4^{2-}} \left[ \frac{*K_3}{(\text{H}^+)} + 1 + (\text{H}^+)/*K_2 + (\text{H}^+)^2 / *K_1 *K_2 \right]$ , where  $m$  refers to aqueous concentrations in molal units,  $(\text{H}^+)$  is the activity of hydronium ions in solution, and  $*K_1$ ,  $*K_2$ ,  $*K_3$  are the conditional (stoichiometric) ionization constants of phosphoric acid. Combining the mineral solubility equilibria, the phosphoric acid dissociation equilibria and the mass balance constraints,  $m_{\text{Ca}/5} = m_{\text{TP}}/3 = m_{\text{HAP}}$  and  $m_{\text{Ca}/5} = m_{\text{TP}}/3 = m_{\text{F}} = m_{\text{HAP}}$ , the total dissolved

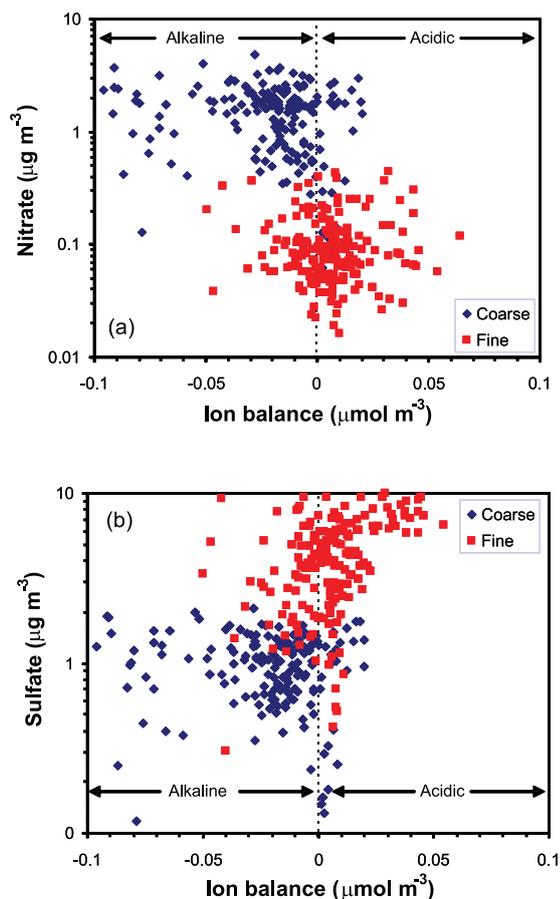
phosphate concentrations in equilibrium with HAP and FAP can be computed.

Representative solubility calculations were performed for HAP and FAP dissolving in pure electrolyte (NaCl) solutions at 25 °C. The conditional ionization constants for phosphoric acid (measured up to 6 *m* NaCl) were taken from Hershey et al. (1989). The Pitzer ion interaction model was used to calculate the activity coefficients (Hershey et al., 1989; Millero and Schreiber, 1982). The latter is applicable up to 3–4 *m* NaCl, typical of electrolyte concentrations calculated using the ISORROPIA-II model predictions for aerosol at ~95 % RH.

### 2.4 Results from Eastern Mediterranean aerosol analysis

Under typical marine conditions (relative humidity 80–95 %), soluble electrolytes in the aerosol cause the uptake of water from the gas phase, forming an aqueous solution that is in contact with the insoluble components of the dust aerosol (Seinfeld and Pandis, 2006). This aerosol solution provides the medium for reacting and dissolving mineral-bound P. A proxy of the acidity of the aerosol aqueous phase is given by the soluble ion balance,  $I_b = 2[\text{SO}_4] + [\text{H}_2\text{PO}_4] + [\text{NO}_3] + [\text{Cl}] - 2[\text{Ca}] - [\text{NH}_4] - [\text{Na}] - 2[\text{Mg}] - [\text{K}]$  (where “[ $X$ ]” represents the measured concentration of species “ $X$ ” expressed in  $\mu\text{moles per m}^3$  air, and  $I_b$  has units of  $\mu\text{greqs per m}^3$  air, “greqs” being “gram-equivalents”). When  $I_b > 0$ , the aerosol tends to be acidic (as excess  $\text{H}^+$  is required in the associated aqueous phase to neutralize the excess anions) and vice versa. The coarse (diameter  $> 1.3 \mu\text{m}$ ) and fine (diameter  $< 1.3 \mu\text{m}$ ) aerosol fractions exhibit distinct acidities (Fig. 1):  $I_b$  (and aerosol acidity) in the fine fraction is largely controlled by sulfate and in the coarse fraction by nitrate. When the aerosol is acidic ( $I_b > 0$ ), nitrate tends to partition to the gas-phase in the form of nitric acid (Meskhidze et al., 2003). When the aerosol is neutral or alkaline ( $I_b < 0$ ), nitrate partitions to the aerosol phase, displacing chloride and carbonate ions (from seasalt and carbonate minerals) to the gas phase (in the form of HCl and  $\text{CO}_2$ , respectively; Meskhidze et al., 2003). As expected, coarse particles contain relatively low sulfate levels, typical of marine background aerosols ( $\sim 0.1\text{--}1 \mu\text{g m}^{-3}$ ; Seinfeld and Pandis, 2006). Sulfate concentrations of fine particles were distinctly higher, indicative of continental pollution ( $\sim 1\text{--}10 \mu\text{g m}^{-3}$ ; Seinfeld and Pandis, 2006). Aerosol nitrate and sulfate concentrations no longer correlate with  $I_b$  when  $I_b < -0.05 \mu\text{greqs m}^{-3}$ , probably because insufficient acid was present to titrate the carbonate buffer of dust.

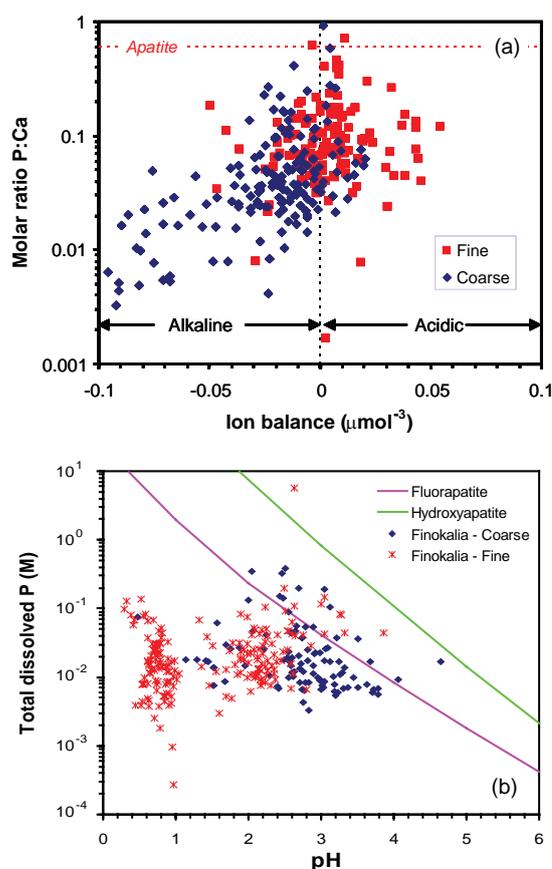
For alkaline particles the measured mass ratio of soluble phosphorous to soluble calcium generally fell well below the P:Ca ratio of the mineral apatite (Fig. 2a). As aerosol acidity increased however, the soluble P:Ca ratios also increased and approached the value expected for the stoichiometric dissolution of apatite. The scatter in the data can in part be attributed



**Fig. 1.** Distributions of key acidic aerosol species in the Finokalia data set. Shown are the concentrations of (a) nitrate and (b) sulfate against the “ion balance”,  $I_b$ , calculated as  $I_b = 2[\text{SO}_4] + [\text{H}_2\text{PO}_4] + [\text{NO}_3] + [\text{Cl}] - 2[\text{Ca}] - [\text{NH}_4] - [\text{Na}] - 2[\text{Mg}] - [\text{K}]$ , where “[X]” represents the concentration of species “X” in the aerosol sample, in  $\text{mol m}^{-3}$  air. Data for coarse and fine aerosol fractions are presented as blue and magenta symbols, respectively.

to chemical (and, hence, pH) variability across particle size and time, which were averaged during aerosol collection (the pH of individual particles depends especially on their carbonate content which is known to be variable; samples collected over several days may also combine particles from somewhat different air masses). Nevertheless, the observed trend between the soluble P:Ca ratio and  $I_b$  is strongly consistent with acid mobilization of phosphate from apatite, the dominant form of P in Saharan dust (see below). Examining the data for periods where fires are at a minimum (October–March) results in a plot very similar to Fig. 2a (not shown); thus biogenic P from biomass burning is not the prime cause for the correlation between acidity and P:Ca.

In itself,  $I_b$  does not directly yield the exact pH level in the aerosol solution nor the amount of water available for dissolution of P from dust. Both can be estimated using ISORROPIA-II (Fountoukis and Nenes, 2007; Sect. 2.2). Using typical levels of relative humidity and



**Fig. 2.** (a) The molar ratio of LIP over Ca as a function of ion balance,  $I_b$ . Fine fraction samples are shown with magenta symbols, coarse particles with blue symbols. Also shown is the P:Ca ratio for the mineral apatite. (b) Predicted concentration of total dissolved P against aerosol pH. Predictions are carried out using the ISORROPIA-II thermodynamic model (Fountoukis and Nenes, 2007), assuming that the aerosol particles are at thermodynamic equilibrium with an ambient relative humidity (RH) of 95 %, a temperature of 298 K, and with aqueous compositions constrained from the observations. Also shown are the solubility lines for the end-member minerals fluorapatite,  $\text{Ca}_5(\text{PO}_4)_3\text{F}$  and hydroxyapatite,  $\text{Ca}_5(\text{PO}_4)_3\text{OH}$ . Repeating the calculations for lower RH decreases pH, so that the aerosols become even more unsaturated with respect to the apatite minerals. For clarity only the solubility lines at ionic strength 3 m are plotted.

observed temperatures, the vast majority of fine aerosols at Finokalia exhibited low pH and were predicted to be highly undersaturated with respect to the most insoluble end-member apatite mineral, crystalline fluorapatite (Fig. 2b). Thus, thermodynamically, soil-derived apatite and iron-bound P are expected to dissolve during aerosol acidification. Variations in relative humidity do not affect this model outcome. In fact, pH could substantially increase above 4 only if the aerosol particles were to reside in clouds and were activated into cloud droplets. Given however that atmospheric

particles reside in subsaturated water vapor (non-cloud) conditions for most of their atmospheric residence time (Seinfeld and Pandis, 2006), the pH predictions and solubility considerations presented in Fig. 2b reflect the most frequently occurring state in the atmosphere.

### 3 Testing P mobilization hypothesis in the laboratory

The Finokalia data and subsequent thermodynamic analysis suggest that acidification plays an important role in producing LIP. We tested the proposed mechanism by exposing Saharan soil and dust samples to pH levels consistent with the aerosol acidities derived from the Finokalia data. Soil samples were collected from two locations known to be major sources of modern Sahara dust: the Tibesti Mountains (hereafter Tibesti), and the Western Sahara (hereafter WS) (Brooks and Legrand, 2000; Prospero et al., 2002; Washington et al., 2009). The samples were wet-sieved to less than 20  $\mu\text{m}$  ( $\text{PM}_{20}$ ). We also generated  $\text{PM}_{10}$  (particulate matter less than 10  $\mu\text{m}$ ) from the re-suspended soils (called Tibesti- $\text{PM}_{10}$  and WS- $\text{PM}_{10}$  hereafter) without wet sieving. Details and chemical compositions of the samples are described in Shi et al. (2011a). In addition, two atmospheric Saharan dust samples were examined; they were collected at Bet Yanai, Israel, after dust storms on 10 May 2004 and on 29 March 2008 (hereafter BY1 and BY2, respectively). Prior to acid processing, the samples were characterised for the amounts of LIP, iron-bound mineral P (Fe-P), calcium-bound mineral P (Ca-P), and organically-bound P (Org-P) using the SEDEX procedure (Ruttenberg, 1992). In order to simulate atmospheric processing, the mineral samples were placed in pH 2 sulphuric acid for 24 h.

#### 3.1 Phosphorus speciation of samples

All samples were characterised for the amount of LIP, iron-bound phosphorus (Fe-P), Ca-bound phosphorus, residual inorganic P, and organic phosphorus (OP) using the SEDEX procedure (Ruttenberg et al., 2009):

- LIP: about 100 mg of dust or soil samples were leached for 2 h in 10 ml of 1 M  $\text{MgCl}_2$  solution adjusted to pH 8 with NaOH. The solutions were then filtered through 0.2  $\mu\text{m}$  pore size filters and the residues were treated with the same  $\text{MgCl}_2$  solution for another 2 h and then filtered. P in both extractions was added to obtain the LIP.
- Fe-P: 0.37 g sodium dithionite was added to the residue of previous extraction; 10 ml of citrate buffer solution (88.41  $\text{g l}^{-1}$  trisodium citrate and 84.32  $\text{g l}^{-1}$  sodium bicarbonate) was then added to each sample. The solutions were shaken for 8 h at 25  $^\circ\text{C}$ . The samples were filtered through a 0.2  $\mu\text{m}$  pore size filter. Then 10 ml  $\text{MgCl}_2$  solution (same as above) was added to the

residues and left for 2 h. The samples were then filtered again through a 0.2  $\mu\text{m}$  pore size filter. The measured P was defined as Fe-P.

- Ca-bound-P: the residues were then treated with 10 ml 1 M sodium acetate solution buffered at pH 4 with acetic acid for 6 h and then filtered. The residues were then leached with  $\text{MgCl}_2$  solutions (same as above) two times. The filtrates were mixed and then diluted for 5 times with double distilled water.
- Residual inorganic P: the residues were then treated with 10 ml 1 M HCl for 16 h and then filtered. The Ca-bound-P and residual inorganic P added together was defined as Ca-P.
- Org-P: The residues were ashed at 550  $^\circ\text{C}$  for 2 h and treated with 10 ml 1 M HCl for 16 h and then filtered. The P was defined as OP.

Ca-bound P and residual inorganic P are defined as Ca-P. Fe-P + Ca-P are defined as total inorganic P. And the total inorganic P plus Org-P is total-P.

In addition to SEDEX, we also measured the seawater leachable P and total inorganic P (TIP) using independent method.

- Sea water leachable P: about 10 mg of Tibesti, WS, BY1, or BY2 samples were put into 15 ml tubes, and leached into seawater for 30 h. The seawater was filtered through a 0.2  $\mu\text{m}$  pore size polycarbonate filter and was added with 50  $\mu\text{l l}^{-1}$  of chloroform. After 30 h, the samples were filtrated through a 0.2  $\mu\text{m}$  pore size Teflon filter. The dissolved phosphorus content of the seawater sample after filtration (and those of the blanks) was analyzed within an hour.
- TIP: about 10 mg of Tibesti, Tibesti- $\text{PM}_{10}$ , WS, or WS- $\text{PM}_{10}$  samples were leached with 10 ml of 1 M HCl. The solution was left at room temperature overnight (16 h) with constant shaking. The solutions were then filtrated through 0.2  $\mu\text{m}$  pore size filters and the filtrates analysed for phosphorus.

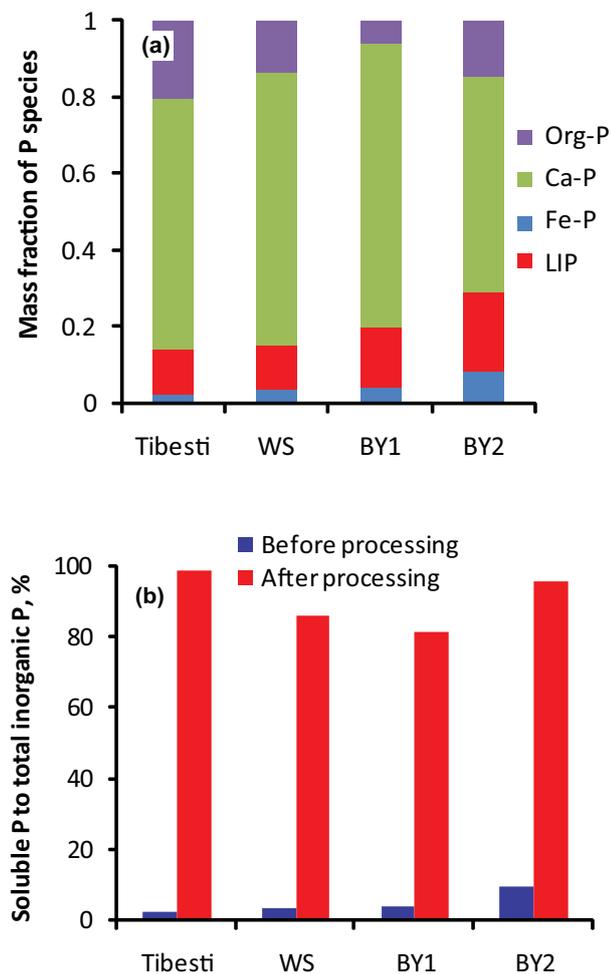
Replicate analysis of Tibesti samples ( $n = 4$ ) resulted in relative standard deviation of 14.2, 9.4, 10.6, 3.0 and 7.2 % respectively for LIP, Fe-P, Ca-bound P, org-P and total-P. The measured P concentration in each extraction solution was corrected for the blanks.

TIP measured with SEDEX and with 1 M HCl only were 17.5 and 17.3  $\mu\text{mol g}^{-1}$  in the Tibesti sample, and 18.6 and 18.5  $\mu\text{mol g}^{-1}$  in the WS sample, respectively. TIP measured by 1 M HCl extractions in the Tibesti (17.3  $\mu\text{mol g}^{-1}$ ) and WS (18.5  $\mu\text{mol g}^{-1}$ ) were almost the same as that in Tibesti- $\text{PM}_{10}$  (17.5  $\mu\text{mol g}^{-1}$ ) and WS- $\text{PM}_{10}$  (18.2  $\mu\text{mol g}^{-1}$ ), respectively, suggesting that the  $\text{PM}_{20}$  is representative of the  $\text{PM}_{10}$ .

**Table 1.** Dust samples and simulated atmospheric processing results. All concentrations are given in  $\mu\text{mol}$  per gram dry weight.

Sample	Seawater LIP	LIP	Fe-P	Ca-P	TIP	Org-P	Total-P	pH 2 LIP
Tibesti	0.3	0.5	2.6	14.4	17.5	4.5	21.9	16.7
WS	0.1	0.7	2.5	15.3	18.4	2.8	21.3	15.8
BY1	0.9	1.3	5.3	24.8	31.3	2.0	33.4	25.4
BY2	4.0	3.2	8.1	22.0	33.3	5.6	38.8	31.8
BYD1*		1.9	5.3	18.0	25.2	3.2	28.4	

\* Data from Eijssink et al. (2000).



**Fig. 3.** (a) Phosphorus speciation in Saharan soil and natural Saharan dust. (b) The change in soluble P to total inorganic P as a result of simulated atmospheric processing at pH 2 for 24 h.

### 3.2 Laboratory simulated atmospheric acid processing and analysis

About 10 mg of Tibesti-PM<sub>20</sub>, Tibesti-PM<sub>10</sub>, WS-PM<sub>20</sub>, WS-PM<sub>10</sub>, BY1 or BY2 were added to 10 ml pH 2 sulfuric acid solution, respectively. The samples were stirred

continuously for 24 h at room temperature. At the end of each experiment, the samples were filtered through 0.2  $\mu\text{m}$  pore size Teflon filters and the solutions taken for subsequent P analysis.

All dissolved phosphorus determinations were carried out by the molybdate-blue method according to Strickland and Parsons (1972) for sea water extractions and Aspila et al. (1976) for all other extractions.

### 3.3 Interpreting laboratory acidification experiments

In agreement with previous studies (Singer et al., 2004; Eijssink et al., 2000), we found that Ca-P was by far the dominant form of P in Saharan soil dust (>50 %) while iron-bound P represents less than 20 % (Table 1). LIP represented only 3–10 % of total inorganic P in Saharan soil samples and dry fallout collected during Sahara dust storms (Fig. 3a). Acid treatment of the samples caused a 10–40 times increase of their soluble P fraction of total inorganic P (Fig. 3b). In the Saharan soil and dust samples, 81–96 % of the total inorganic P was released to solution following the acid treatment. Mass balance considerations show that this soluble P was derived from both apatite and iron-bound P. The experimental results overall showed that significant fractions of P present in Saharan dust can be solubilised by acidification in the atmosphere. Upon deposition of the aerosols at the sea surface, the solubilised P will remain in solution and be available for direct uptake by phytoplankton (Herut et al., 2005). This is in contrast to soluble iron which precipitates as Fe oxyhydroxide nanoparticles under the alkaline and oxygenated conditions in surface waters (Shi et al., 2009).

## 4 Conclusions and implications

In recent years, it has become apparent that P (co-)limitation in the ocean is more widespread than previously thought (Elser et al., 2007; Paytan et al., 2007). In view of this, our results suggest that variations in acid gas emissions by natural and anthropogenic sources may play an important role in regulating marine primary productivity and autotroph nutrient limitation through their effect on the fraction of soluble P in mineral aerosols. While atmospheric pollution is widely

acknowledged as a major source of new nitrogen (N) to the global ocean (Duce et al., 2008) we propose that anthropogenic SO<sub>2</sub> and NO<sub>x</sub> emissions may also increase the input of bioavailable P. Because primary production in marine environments often exhibits a positive synergistic response to simultaneous enrichments in N and P (Elser et al., 2007), SO<sub>2</sub> and NO<sub>x</sub> emissions have the potential to significantly enhance marine primary production, particularly in offshore areas of the ocean where large amounts of mineral aerosol interact with polluted air masses, such as in the EMS, the NW Atlantic or the western Pacific. Similar processes are known to affect the fraction of bioavailable Fe in mineral dust (e.g. Shi et al., 2009, 2011b). The interaction between these two elements in atmospheric aerosols will no doubt be the subject of future research. In the extreme case that 100 % of atmospheric P deposition were bioavailable it could, assuming P limitation, account for a yearly global oceanic uptake of CO<sub>2</sub> of 264 T moles, that is, more than three times the current CO<sub>2</sub> emissions from fossil fuel burning (Mahowald et al., 2008). As shown here, even if total P deposition remains constant, the soluble fraction can significantly increase as a result of aerosol acidification (Fig. 3b).

Soluble P production by aerosol acidification also helps explain the proposed link between episodes of massive volcanism and global ocean anoxia in the geological past (Adams et al., 2010). On geological time scales, ocean productivity is P limited because the marine biosphere can accommodate any long-term nitrogen deficiency by increasing nitrogen fixation (Paytan et al., 2007; Tyrell, 1999). A massive release of acidic gases by volcanic activity would then dramatically increase the acid processing of aerosols in the atmosphere (as well as enhance chemical weathering on land). The resulting increase in bioavailable P supply to the oceans would trigger a major increase in primary productivity, which in turn would enhance the supply of organic carbon to the deep ocean eventually resulting in widespread ocean anoxia with its distinctive chemical and isotopic signatures (Adams et al., 2010).

On shorter time scales, the potential also exists for a phytoplankton-mediated feedback between aerosol acidification and ocean productivity. Under conditions of nutrient stress, marine phytoplankton increase emissions of dimethyl sulfide (DMS) from the surface ocean (Sunda et al., 2002). The DMS then converts to sulphuric and methane-sulfonic acid in the atmosphere (Meskhidze et al., 2003; Barnes et al., 2006). Based on the proposed hypothesis, this would in turn increase atmospheric deposition of bioavailable P, hence alleviating the initial nutrient stress. It would thus appear that atmospheric processes could play a far greater role in the biogeochemical cycling of phosphorus than is currently acknowledged.

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